DIFFERENT REACTIONS with ACIDS

NEUTRALISATION REACTIONS

- This is when *an **acid** reacts with a **base*** to produce a **neutral** salt, and **water**.
- So the acid is ***neutralised*** by the base.

Acid + Base → Salt + Water

- The **H** from the acid reacts with the **OH** from the base. This is where the **H_2O** comes from.

HCl + Na**OH** → NaCl + **H_2O**

HydroChloric Acid + Sodium HydrOxide → Table Salt + Water

See other examples of Neutralisation Reactions on Page 39.

ACID + METAL → SALT + HYDROGEN

$2HCl + Mg \rightarrow MgCl_2 + H_2$

ACID + METAL-*HYDRO*XIDE → SALT + WATER

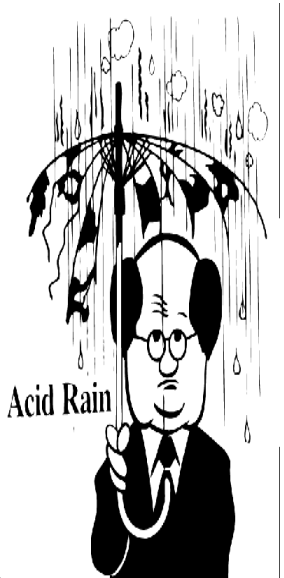
$HCl + NaOH \rightarrow NaCl + H_2O$

ACID + METAL-OXIDE → SALT + WATER

$HCl + MgO \rightarrow MgCl_2 + H_2O$

ACID + METAL-CARBONATE → SALT + CO_2 + WATER

$HCl + CaCO_3 \rightarrow CaCl_2 + CO_2 + H_2O$



ACID RAIN

O_2, CO_2 and N-Oxides + H_2O → Acid Rain

*(See details of **Acid Rain** on Page 41.)*